

## Local Coordination Structures of $\gamma$ -Al<sub>2</sub>O<sub>3</sub>-Supported Ir Dimer for Surface-Assisted Transfer Hydrogenation Catalysis

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### Introduction

A novel oxide-supported Ir dimer, which was found to be active for transfer hydrogenation of aromatic ketones, was prepared on a  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> surface. Detail characterization of the supported Ir dimer revealed the surface-assisted transfer hydrogenation catalysis on the supported Ir dimer catalyst. Local coordination structures of the supported Ir dimers were investigated by Ir L<sub>III</sub>-edge XAFS.

### Experimental

An Ir dimer [Ir<sub>2</sub>{ $\eta^5$ -C<sub>5</sub>(CH<sub>3</sub>)<sub>5</sub>}<sub>2</sub>( $\mu$ -CH<sub>2</sub>)<sub>2</sub>] (**Ir**<sub>2</sub>) was impregnated on several oxide supports ( $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, SiO<sub>2</sub>, and MgO) and reacted with surface hydroxyl groups on these surfaces at 473 K for 2 h (**Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, **Ir**<sub>2</sub>/SiO<sub>2</sub>, and **Ir**<sub>2</sub>/MgO). Hydride intermediate Ir dimer was prepared by the reaction of the supported Ir dimer on  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> and isopropanol at 343 K (**Ir**<sub>2</sub>-H<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>), and subsequent reaction with acetophenone was performed in toluene at 343 K (**Ir**<sub>2</sub>-Rxn1/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>).

Ru K-edge XAFS was measured at BL9C and BL12C stations with Si(111) double crystal monochromator. Ionization chambers filled with pure N<sub>2</sub> and Ar gases were used to monitor the incident and transmitted X-rays, respectively. EXAFS spectra were analyzed using REX2000, Ver. 2.3.3, Rigaku. Curve-fitting analysis was carried out in *k*-space and *k*<sup>3</sup>-weighted EXAFS oscillations were Fourier transformed into *R*-space. The fitting parameters were coordination number (CN), interatomic distance (*R*), Debye–Waller factor ( $\sigma^2$ ; mean-square-displacement), and the correction-of-edge energy ( $\Delta E_0$ ). Phase shifts and backscattering amplitudes for Ir–Ir and Ir–C (C<sub>5</sub>(CH<sub>3</sub>)<sub>5</sub>) were calculated by using FEFF8 code.

### Results and Discussion

The local coordination structure of **Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> was investigated by the curve-fitting analysis of Ir L<sub>III</sub>-edge EXAFS data using the results of quantitative analysis of the coordinating ligands (Table 1). The precursor **Ir**<sub>2</sub>, whose coordination structure was determined by X-ray diffraction, was successfully fitted and the EXAFS oscillation of **Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> was fitted well with four contributions: Ir–Ir at 0.269 ± 0.001 nm (CN = 1.2 ± 0.2), Ir–C (C<sub>5</sub>(CH<sub>3</sub>)<sub>5</sub>) at 0.245 ± 0.001 nm (CN = 5.0 ± 0.7), Ir–O at 0.212 ± 0.001 nm (CN = 3.2 ± 0.2), and Ir···Al at 0.319 ± 0.001 nm (CN = 1.9 ± 0.2). The Ir–Ir bond distance at 0.269 ± 0.001 nm indicates that the Ir=Ir bond of **Ir**<sub>2</sub> was stretched to a single Ir–Ir bond on the  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> surface. The CN of Ir···Al (1.9 ± 0.2) at 0.319 ± 0.001

nm suggested that there were two O(Al) species coordinated to the Ir center. The CN of Ir–O/C was estimated to be 3.2 ± 0.2, where the oxygen and carbon contributions cannot be discriminated by EXAFS curve fitting. The quantitative analysis by GC indicated that one  $\mu$ -CH<sub>2</sub> ligand remained on the Ir dimer (the CN of Ir–C ( $\mu$ -CH<sub>2</sub>) = 1), and hence the CN of Ir–O was estimated to be 2, which was similar to that of Ir···Al (1.9 ± 0.2).

Ir L<sub>III</sub>-edge EXAFS revealed that **Ir**<sub>2</sub>-H<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> maintained its original dimeric structure with the Ir–Ir bond at 0.267 ± 0.001 nm (CN = 0.9 ± 0.1). It is to be noted that the CN of Ir···Al dropped from 1.9 ± 0.2 at 0.319 ± 0.001 nm to 1.2 ± 0.2 at 0.321 ± 0.001 nm. The CN of Ir–O/C was also reduced from 3.2 ± 0.2 (0.212 ± 0.001 nm) to 2.3 ± 0.2 (0.212 ± 0.001 nm), indicating that the CN of Ir–O(Al) was reduced from 2 to 1 under the consideration of the Ir–C ( $\mu$ -CH<sub>2</sub>) contribution (CN = 1). These results suggest that the bridged Ir–(OAl)<sub>2</sub>–Ir bonds in the **Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> became a monodentate Ir–O(Al) bond in each Ir center to form **Ir**<sub>2</sub>-H<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>. It was proposed that the interfacial bond transformation between the Ir dimer (**Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>) and the  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> surface promoted the formation of the dihydride Ir complex on  $\gamma$ -Al<sub>2</sub>O<sub>3</sub> as an intermediate in the transfer hydrogenation.

Table 1 Structural Parameters Determined by Curve-Fitting Analysis of Ir L<sub>III</sub>-edge EXAFS for **Ir**<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, **Ir**<sub>2</sub>-H<sub>2</sub>/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>, and **Ir**<sub>2</sub>-Rxn1/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub>

Shell	CN	Distance /nm	$\Delta E_0$ /eV	$\sigma^2$ /10 <sup>-5</sup> nm <sup>2</sup>
<b>Ir</b> <sub>2</sub> / $\gamma$ -Al <sub>2</sub> O <sub>3</sub> ( <i>R</i> <sub>f</sub> = 0.15%)				
Ir–Ir	1.2 ± 0.2	0.269 ± 0.001	-1	6
Ir–C (C <sub>5</sub> (CH <sub>3</sub> ) <sub>5</sub> )	5.0 ± 0.7	0.245 ± 0.001	-6	9
Ir–O	3.2 ± 0.2	0.212 ± 0.001	17	5
Ir···Al	1.9 ± 0.2	0.319 ± 0.001	20	3
<b>Ir</b> <sub>2</sub> -H <sub>2</sub> / $\gamma$ -Al <sub>2</sub> O <sub>3</sub> ( <i>R</i> <sub>f</sub> = 0.51%)				
Ir–Ir	0.9 ± 0.1	0.267 ± 0.001	-2	4
Ir–C (C <sub>5</sub> (CH <sub>3</sub> ) <sub>5</sub> )	4.9 ± 0.7	0.246 ± 0.001	-10	8
Ir–O	2.3 ± 0.2	0.212 ± 0.001	17	4
Ir···Al	1.2 ± 0.2	0.321 ± 0.001	21	4
<b>Ir</b> <sub>2</sub> -Rxn1/ $\gamma$ -Al <sub>2</sub> O <sub>3</sub> ( <i>R</i> <sub>f</sub> = 0.13%)				
Ir–Ir	1.1 ± 0.1	0.268 ± 0.001	0	5
Ir–C (C <sub>5</sub> (CH <sub>3</sub> ) <sub>5</sub> )	4.9 ± 0.7	0.242 ± 0.001	-9	8
Ir–O	3.2 ± 0.2	0.213 ± 0.001	16	6
Ir···Al	2.2 ± 0.3	0.319 ± 0.001	20	5

Measured at 20 K. *k* = 30 – 150 nm<sup>-1</sup>, *R* = 0.14 – 0.31 nm.

### Reference

[1] M. Tada, S. Muratsugu, M. Kinoshita, T. Sasaki, Y. Iwasawa, *J. Am. Chem. Soc.* **2010**, *132*, 713.

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